

ME 3007
HW 2 Solutions

1.

From distinguishability section:

$$W_{\text{fer}} = \prod_{i=0}^3 \frac{g_i!}{N_{a,i}! (g_i - N_{a,i})!}$$

Because these are fermions, can be distributed



Since particles are indisting, and g is the same for all energy levels, these are interchangeable

For $g=2$:

$$W = \prod_{i=0}^3 \frac{2!}{N_{a,i}! (2 - N_{a,i})!} = \left(\frac{2!}{2!(2-2)!} \cdot \frac{2!}{1!(2-1)!} \cdot \frac{2!}{1!(2-1)!} \cdot \frac{2!}{1!(2-1)!} \right)$$

3 identical

$$W = 1 \cdot 2 \cdot 2 \cdot 2 = \underline{8}$$

For $g=5$:

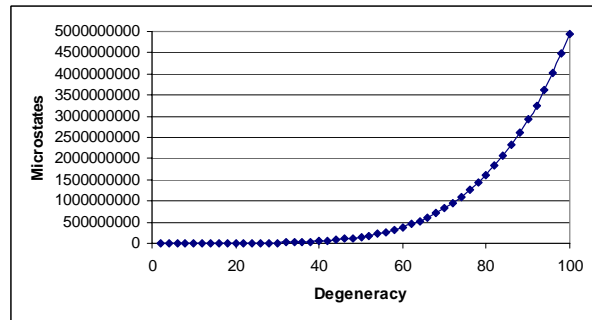
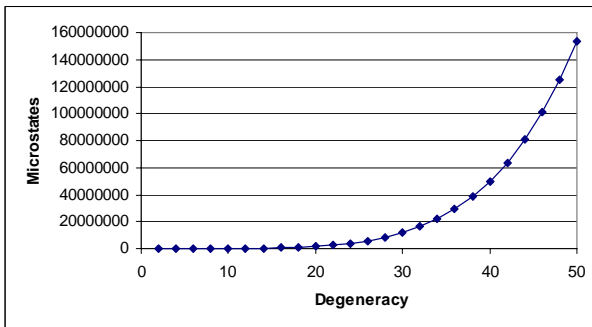
$$W = \left(\frac{5!}{2!(5-2)!} \cdot 3 \left(\frac{5!}{1!(5-1)!} \right)^3 \right) = 1250 \cdot 3 \cdot \frac{5!}{4!}$$

For $g=10$:

$$W = \left(\frac{10!}{2!(10-2)!} \cdot 3 \left(\frac{10!}{1!(10-1)!} \right)^3 \right) = 45,000 \cdot 3 \cdot \frac{10!}{9!} = 11$$

For $g=50$:

$$W = \left(\frac{50!}{2!(50-2)!} \cdot 3 \left(\frac{50!}{1!(50-1)!} \right)^3 \right) = 153,125,000 \cdot 3 \cdot \frac{50!}{49!}$$



2.

Partition Function: $\ln Z = a T^4 V$. I did not specify which ensemble theory to use, so any of canonical, microcanonical, or grand canonical is acceptable. However, since Z is a function of T and V , you might deduce that you should use the canonical ensemble.

For microcanonical:

$$U = \frac{Nk_B T^2}{Z} \left(\frac{\partial Z}{\partial T} \right)_{N,V} = \frac{Nk_B T^2}{\exp(aT^4 V)} (4aT^3 V \exp(aT^4 V)) = \underline{\underline{4Nk_B aVT^5}}$$

$$S = \frac{U}{T} + k_B N \ln Z = 4Nk_B aVT^4 + Nk_B aVT^4 = \underline{\underline{5Nk_B aVT^4}}$$

(Note: This assumes a distinguishable system. If it is indistinguishable, $S = 5Nk_B aVT^4 + Nk_B \dots$)

$$P = k_B T \left(\frac{\partial \ln Z}{\partial V} \right)_{N,U} = k_B T a T^4 = \underline{\underline{ak_B T^5}}$$

(However, here I've had to assume that constant $U = \text{constant } T$, which is not necessarily true.)

For canonical:

$$U = k_B T^2 \left(\frac{\partial \ln Z}{\partial T} \right)_{N,V} = k_B T^2 4aT^3 V = \underline{\underline{4k_B aVT^5}}$$

$$S = k_B \ln Z + \frac{U}{T} = k_B aVT^4 + 4k_B aVT^4 = \underline{\underline{5k_B aVT^4}}$$

$$P = k_B T \left(\frac{\partial \ln Z}{\partial V} \right)_{N,T} = k_B T a T^4 = \underline{\underline{ak_B T^5}}$$

3.

$$\bar{V} = 1020 \text{ m/s}, T = 27^\circ\text{C}, \text{He}, u = ?$$

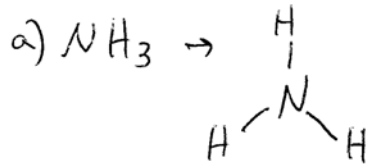
Use the NIST Webbook (or another source) to find that the mass of a He atom is $6.65 \times 10^{-27} \text{ kg}$. The internal energy is due only to the kinetic energy, so

$$e_k = \frac{1}{2} m \bar{V}^2 = 3.46 \times 10^{-21} \text{ J}$$

Avogadro's # gives us the energy per kmol:

$$u = N_A e_k = (6.02 \times 10^{26}) (3.46 \times 10^{-21}) = \boxed{2.082 \times 10^6 \frac{\text{J}}{\text{kmol}} \text{ or } 2082 \frac{\text{kJ}}{\text{kmol}}}$$

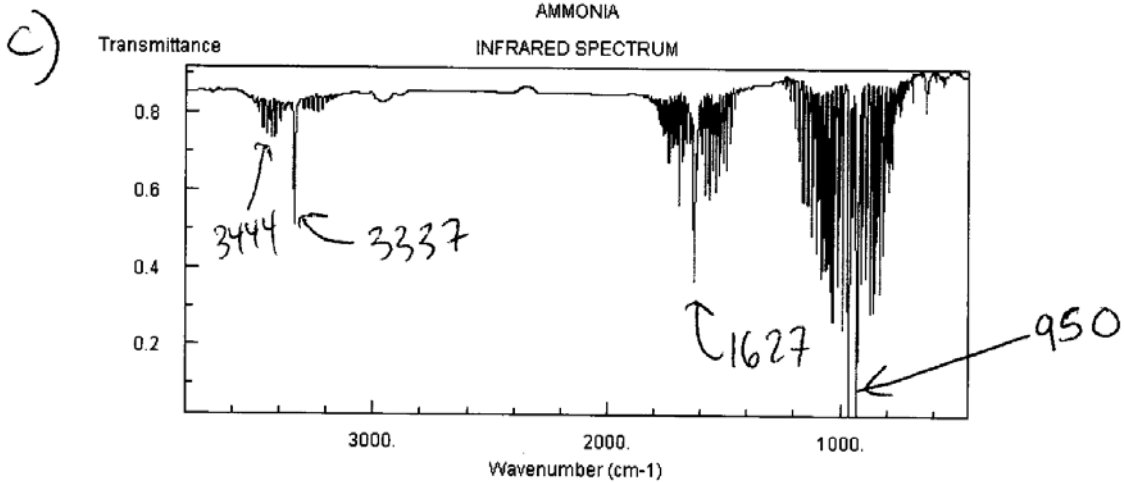
4.



b) 4 molecules \rightarrow 12 DoF

- 3 Translational
- 3 Rotational
- 6 Vibrational

c) You can guess at the peaks from the spectroscopic data, or see in the "Energy Levels" section that they have already been found for you.



Temperature		300	300	600	600	2000	2000
Wave number	Degeneracy	u	E(u)	u	E(u)	u	E(u)
950	1	4.55617	0.222665	2.27808	0.66023	0.6834	0.96197
1627	2	7.80303	0.049785	3.90152	0.64095	1.1705	1.7865
3337	1	16.0041	2.87E-05	8.00207	0.02145	2.4006	0.63186
3444	2	16.5173	3.66E-05	8.25865	0.03535	2.4776	1.22812
Total DoF	6		0.272515		1.35798		4.60845

$$u = 1.4387894 \times \frac{\text{Wave \#}}{T}$$

$$E(u) = \frac{u^2 e^u}{(e^u - 1)^2} \times \text{degen.}$$

Sums

$$C_p = 4R + R \times \text{Sum from Spectrographic data}, R = 8.314$$

$$C_p = 35.52 @ 300 K, 44.546 @ 600 K, 71.57 @ 2000 K$$

d)

	< 1400	> 1400
A	19.99563	52.02427
B	49.77119	18.48801
C	-15.37599	-3.765128
D	1.921168	0.248541
E	0.189174	-12.45799

$$C_p = A + BT + CT^2 + DT^3 + \frac{E}{T^2}$$

$$C_p = 35.7 @ 300K, 45.26 @ 600K, 72.8 @ 2000K$$

∴ excellent agreement